

CH APPROVAL S.1.

Compounds of benzene with halogen-substituted methanes. II. A. P. Kapustinskii and S. I. Drakin (D. I. Mendeleev Chem. Technol. Inst., Moscow). Izv. Akad. Nauk S.S.R., Otdel. Khim. Nauk 1950, 233 0; cf. C.A. 42, 1802c.—Measurements of the vapor pressure, p (in mm. Hg), of C_6H_5Cl over the mol. compd. ClC_6H_5Cl (I) between 9.0 and 25.0° can be represented by the empirical formula $\log p = -(2578/T) + 10.82$, giving for the heat of dissociation of I 12.10 kcal./mole, and (with the known heat of vaporization of C_6H_6 , 8.17 kcal./mole) for the standard heat of formation ΔH°_f of $1,393$ kcal./mole, of the same order of magnitude as ΔH°_f of crystalloborates. I is an instance of a mol. compd. between 2 mols possessing no dipole moment. The conjecture that the formation of a mol. compd. in this case is due to interaction of the π bonds of C_6H_6 with the halogen atom is borne out by the observation that the melting diagram of the binary system ClC_6H_5 -cyclohexane (no unsatd. bonds) shows a continuous series of solid solns. and no sign of compd. formation. This is in keeping with an analogous finding of Timmermans (C.A. 23, 3135) and of Kennard and McEusker (C. 4, 43, 1637c) for CCl_4 -cyclohexane. — N. T.

DRAKIN, S. I.

"Thermochemical Properties of Water Ions." Sub 12 Dec 51, Moscow
Order of Lenin State U imeni M. V. Lomonosov. Cand. Chem. Sci.

Dissertations presented for science and engineering degrees in
Moscow during 1951.

SO: Sum. No. 480, 9 May 55

DRAKIN, S. I.

USSR/Chemistry - Heats of Solution

Apr 52

"Microcalorimeter For Determining Heats of Solution
With a Precision Up To 0.001 Cal," A. F.
Kapustinsky, S. I. Drakin, Chem-Technol Inst
Inenii D. I. Mendeleyev, Moscow

"Zhur Fiz Khim" Vol XXVI, No 4, pp 581-585

A newly designed calorimeter that has a sensitivity
up to 0.00010 and 0.001 cal permits detn of heats
of soln at practically infinite dilns. The values
obtained with its aid for a number of salts agree
with published data. The app is comparatively

217731

simple and permits detn of heats of soln of
difficulty accessible substances, because only
10-30 mg are needed.

217731

KAPUSTINSKIY, A.F.; DRAKIN, S.I.

Thermochemistry of complex compounds. Report no.2: Microcalorimetric determination of the heat of solution in water of ammonium nitroso-pentachlororuthenate. Izv.Sekt.plat.i blag.met. no.27:160-163 '52.
(MLRA 7:5)

1. Institut obshchey i neorganicheskoy khimii im. N.S.Kurnakova Akademii nauk SSSR. 2. Moskovskiy khimiko-tehnologicheskiy institut im. D.I.Mendelejeva. (Thermochemistry) (Compounds, Complex) (Ruthenates)

"APPROVED FOR RELEASE: Friday, July 28, 2000 CIA-RDP86-00513R0004111200

APPROVED FOR RELEASE: Friday, July 28, 2000 CIA-RDP86-00513R00041112001

REMARKS, S. T.

USSR/Chemistry - Borillium Salts, cobalt
Salts

Apr 53

"An adiabatic Calorimeter for Determining the Heat Capacities of Solutions. Investigation of salts of Co²⁺, Ni²⁺, Be²⁺, and Mg²⁺ A. F. Kapustinskii, V. V. Tukachevskii, S. I. Drakin, Moscow Chem.-Technol. Inst in D. I. Mendeleev

Zhur fiz khim, vol 27, No 4, pp 534-535

Designed an adiabatic calorimeter which permits determination of salts with a precision of ± 0.03%. Investigated especially the heat capacities of salts of BeCl₂, NiCl₂, BeSO₄, and MgCl₂ in the concn range 0.5-2 mol salt per 100 g of water. Confirmed that there is a linear dependence of apparent solar heat capacities on the square root of metal concn. Carried out extrapolation of data to infinite diln and found values of γ for the salts mentioned.

2/0116

"APPROVED FOR RELEASE: Friday, July 28, 2000 CIA-RDP86-00513R0004111200

APPROVED FOR RELEASE: Friday, July 28, 2000 CIA-RDP86-00513R00041112001

DRAKIN, S. I.

USSR/Chemistry - Alloys

Oct 53

"Transfer and Distribution of Components of Metal Alloys in an Electric Field," S. I. Drakin, Chem-Technol Inst im D. I. Mendeleyev

Zhur Fiz Khim, Vol 27, No 10, pp 1586-91.

In an theoretical discussion, gives a new explanation of phenomena which occur in alloys when a direct current is passed. Derives relationships which describe the equilibrium and kinetics of electrodiffusion. Calcd the differences bet

272T17

mean ion charges in the alloys Hg-Na, Hg-Ca, Hg-Li, Hg-K, Hg-Au, Hg-Cd, Hg-Ag, Pd-H, and Fe-C.

U S S R .

✓ Thermodynamic properties of ions in aqueous solution
S. I. Drakin and N. M. Yakubovich
USSR Central Inst. of Metallurgy
1959-60, Vol. 4, No. 2, p. 101-106
The entropy of hydration of a cation in an aqueous soln. is related to its free energy of hydration by the equation $\Delta F = 2.67 \times 10^3 \text{ cal/mole}$, where ΔF and ΔS are expressed in entropy units mole and cal/mole-
deg., resp., and v is the valence of the ion.

USSR/Chemistry - Physical chemistry

Card 1/1 Pub. 147 - 25/27

Authors : Drakin, S. I.

Title : Geometrical model of an ion solution

Periodical : Zhur. fiz. khim. 28/9, 1698-1700, Sep 1954

Abstract : The use of a geometrical model (ball bearings 2mm in diameter) to determine the effect of ions on the structure of solutions is discussed. Results of such a model investigation are described. The use of small balls is explained by the fact that the intensity field of mono-atomic ions is perfectly symmetrical and the form of these ions is spherical. Five references: 3-USSR; 1-Hungarian and 1-USA (1934-1953). Graph.

Institution : The D. I. Mendeleyev Chemical-Technological Institute, Moscow

Submitted : April 29, 1954

"APPROVED FOR RELEASE: Friday, July 28, 2000 CIA-RDP86-00513R0004111200

APPROVED FOR RELEASE: Friday, July 28, 2000 CIA-RDP86-00513R00041112001

DURKIN, S. I.

"Electrodiffusion in Liquid Potassium-Sodium Alloys"

A Report given at a Scientific-Technical Conference at the
Moscow Chemicotechological Inst. in D.I. Mendeleev, 22-29 Feb 56.

Sum 1137, 28 Nov 56

DRAKIN, S. I.

Category: USSR / Physical Chemistry
Thermodynamics. Thermochemistry. Equilibrium. Physico-
chemical analysis. Phase transitions.

B-8

Abs Jour: Referat Zhur-Khimiya, No 9, 1957, 29950

Author : Drakin S. I.

Inst : Moscow Chemico-Technological Institute

Title : Thermal Capacity of Aqueous Solutions of Aluminum Sulfate and
Potassium Fluoride

Orig Pub: Tr. Mosk. khim.-tekhnol. in-ta, 1956, No 22, 79-83

Abstract: Determinations, at 25°, of thermal capacity of aqueous solutions
of Al₂(SO₄)₃ (I) in the concentration range from 0.1 to 1 M, and
of KF from 0.3 to 9.3 M, and calculations of the apparent thermal
capacity (T) of salts in solution. Determination procedure was
described previously (RZhKhim, 1954, 17799). Accuracy of deter-
mination ± 0.05% in the case of KF solutions, and of ± 0.07% in
the case of dilute (below 10%) and of ± 0.2% in the case of more
concentrated solutions of I. Dependence of T of I on \sqrt{m}

C

Card : 1/2

-70-

DRAKIN, S.I.

DRAKIN, S.I.; MAL'TSEV, A.K.

Electrodiffusion in potassium-sodium alloys [with summary in
English] Zhur.fiz.khim. 31 no.9:2036-2041 S '57. (MIRA 11:1)

1. Moskovskiy khimiko-tehnologicheskiy institut im. D.I. Mendeleyeva.
(Potassium-sodium alloys) (Diffusion) (Electrochemistry)

DRAKIN, S. I. and YATSIMIRSKIY, K. B. (Ivanova)

"The Entropy of Ion-Solvation,"

Report presented at Conference on the Effect of Solvents on the Properties of Electrolytes, Khar'kov, 14-15 Oct '57.

Zhurnal Fizicheskoy Khimii, 1958, Vol 32, Nr 4, pp 960-962.

5(4)

AUTHORS: Drakin, S. I., Mikhaylov, V. A. SOV/76-33-7-15/40

TITLE: Calculation of the Entropy of the Hydration of Cations

PERIODICAL: Zhurnal fizicheskoy khimii, 1959, Vol 33, Nr 7, pp 1544-1550
(USSR)

ABSTRACT: In a previous paper (Ref 1), a method of calculating the hydration energy of cations ΔZ_{gas} was described on the basis of a thermodynamic cycle, and a corresponding equation (1) was obtained. The latter permits calculation of the hydration entropy of the ions ΔS_{gas} if the application of this method is restricted to a particular case. In the present case, the components of equation (1) were differentiated with respect to temperature, and corresponding equations were obtained which permit calculation of the hydration entropy (HE) of cations that do not possess the electronic configuration of inert gases. According to the resultant equations, the (HE) of 34 cations was determined and compared with experimental data (Table) which were in good agreement. The majority of mono- and diatomic ions deviate only by 7 entropy units at most. The above method of calculating (HE) permits explanation of the linear dependence of

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Calculation of the Entropy of the Hydration of Cations SOV/76-33-7-15/40

(HE) on z/r_i (z = charge, r_i = ionic radius) observed by (Ref 10), which could not be explained by Born's equation. From the resulting values of hydration energy and entropy the authors calculated the hydration heats of the above cations (Table), which were in good agreement with experimental data. There are 2 figures, 1 table, and 13 references, 7 of which are Soviet.

ASSOCIATION: Khimiko-tekhnologicheskiy institut im. D. I. Mendeleeva, Moskva
(Institute of Chemical Technology imeni D. I. Mendeleyev, Moscow)

SUBMITTED: December 31, 1957

Card 2/2

AUTHORS: Kholpanov, L. P., Drakin, S. I. 8/193/60/003/01/003/058
TITLE: Derivation and Solution of the Differential Calculus Describing
the Transport of Substance in Electric Diffusion of Metal Alloys
PERIODICAL: Izvestiya vysshikh uchebnykh zavedeniy. Khimiya i khimicheskaya
tekhnologiya, 1960, Vol 3, Nr 1, pp 14-19 (USSR)

TEXT: The authors derive an equation (29) calculating the concentration (or impoverishment) of an alloy if 2 parameters K and D characterizing electric diffusion are known. The transport of substance during electric diffusion can be considered the final result of two processes of opposite direction: a) the migration of the alloy components under the action of electric current producing a concentration gradient, b) back-diffusion released by the above gradient. With a sufficiently long current passage, a steady state develops: the effect of the above factors is counterbalanced, and the transport of substance ceases. Equation (29) derived by the authors characterizes the transport in different points of the alloy at any point of time. As the solution of this problem meets with considerable mathematical difficulties, the authors introduced the following restrictions: only diluted metal solutions are investigated; the electric diffusion is studied in an alloy contained in a tube of a constant cross section q; the tube lies in the axis x. After extensive mathematical computations starting from equation (1), the authors arrive

Card 1/2

Derivation and Solution of the Differential Calculus
Describing the Transport of Substance in Electric Diffusion of Metal Alloys

S/153/60/003/01/003/058
D011/B005

at equation (29). The parameters K and D required for the use of equation (29) are determined as follows: K is found from data for steady distribution (Refs 1,2). The ratio c/c_0 (c = concentration of the dissolved metal in equation (1), c_0 = this concentration in all points of the alloy at the beginning of the electric diffusion process: $c(x,0) = c_0$) is measured in any section of the tube shortly after the beginning of the current passage. Thus, the value D can be determined by means of equation (29). It should not be forgotten that equation (29) was derived for an infinitely long tube whereas in practice only finitely long tubes are available (Fig 1). Finally, the authors complain that very few data are found in publications to check equation (29). Figures 2 and 3 are taken from reference 7. They show that the value K, calculated according to reference 7, seems to be much too low. B. Baranovskiy is mentioned in the paper. There are 3 figures and 8 references, 4 of which are Soviet.

ASSOCIATION: Moskovskiy khimiko-tehnologicheskiy institut im. D. I. Mendeleyeva;
Kafedra fizicheskoy khimii (Moscow Institute of Chemical Technology
imeni D. I. Mendeleyev; Chair of Physical Chemistry)

SUBMITTED: April 27, 1959

Card 2/2

(C)

MIKHAYLOV, V.A.; BRAKIN, S.I.

Mechanism of the solvation of ions. Izv.Sib.otd.AN SSSR no.6:44-
52 '60. (MIRA 13:9)

1. Institut neorganicheskoy khimii Sibirskogo otdeleniya AN SSSR
i Moskovskiy khimiko-tehnologicheskiy institut im. D.I. Mendeleyeva.
(Solvation)

DRAKIN, S.I.; GOLUBKOVA, Yu.K.

Electrodiffusion in dilute solutions of lead and mercury in metallic potassium. Zhur. fiz. khim. 34 no.4:866-871 Ap '60.(MIRA 14|5)

1. Khimiko-tekhnologicheskiy institut imeni D.I.Mendeleyeva, Moskva.
(Diffusion) (Lead) (Mercury) (Potassium)

54700

24022
S/0/6/61/035/005/005/005
B1C1, R16

AUTHORS: Drekin, S. I., Sergeyeva, T. N., and Rusekova, V. N. (Moscow)

TYPE: Study of electrodiffusion in the alloys K-Tl, Na-Eg, Na-Pt,
and Na-Cd

PUBLICATION: Zhurnal fizicheskoy khimii, v. 32, no. 6, 1958, 115-1152

TEXT: In previous papers (Rif. 1, 2; Zh. fiz. khimii 31, 2056, 1957; ibid., 31, No. 1060) it has been found that all alkali metals are able to enter into studying electrodiffusion. In order to clarify the mechanism of this effect, the authors studied electrodiffusion in the alloys: Na-Eg containing 0.6 and 1.1% Eg; K-Tl containing 0.19% Tl; Na-Tl containing 0.04% Pt and Na-Cd containing 0.01% Cd. The experimental method is indicated in Ref. 1. The experiments were made in a special glass tube (by means of resistance heating) in which contained the alloy 2, and in the resistance maintained 3 was heated. The alloys were packed into the tube by a vacuum pump at 110°C, and then a current of 100 mA. After 5 hours (at the end of the tube were cooled) measurements were made in a thermocstat at 110°C. The difference $\Delta\varphi$, volts , was measured.

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24022

S/076/1/05/005/005/008
B101/B210

Study of electrodiffusion in the alkalies.

individual sections of the tube was measured. At 110° the authors assumed φ to be $16.15 \cdot 10^{-5}$ for pure K and $0.96 \cdot 10^{-5}$ for pure Na. Control tests proved that no change of φ occurred in the alkali metals. The following equation is given for the steady state of the diffusion: $\ln(\varphi_1/\varphi_2) = KAE_{1,2}/T + C$, where φ_1 and φ_2 are the concentrations of the dissolved metal in points 1 and 2; $E_{1,2}$ is the voltage drop between points 1 and 2; T - absolute temperature; C denotes a constant with the authors term "coefficient of electrodiffusion". The measurements yielded the following values for C: $2.14 \cdot 10^5$ for Na-Hg containing 0.5% of Hg, $0.90 \cdot 10^5$ deg/v for Na-Hg containing 1.1% of Hg; and $2.27 \cdot 10^5$ deg/v for K-Tl. The values for Na-Pb and Na-Cd could not be determined accurately because the concentration of the metal dissolved was very low. The following mean values are given: for Na-Pb ($2.6 \pm 1.5 \cdot 10^5$ deg/v); for Na-Cd ($1.5 \pm 0.5 \cdot 10^5$ deg/v). The authors discuss the equation by S. I. Drakin (Zh. fiz. khimii, 27, 1566, 1953): $C = F(\psi'' - \psi')/R$ (2), (ψ', ψ'') are the effective ionic charges of the metal dissolved and of the metal serving as

Card 2/4

Study of electrodiffusion in the alloys...

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S/076/61/035/005/005/008
B101/B218

solvent, respectively; F = Faraday number; R - gas constant). This equation which has been criticized already by B. Baranowski (Zh. fiz. khimii, 28, 1676, 1954), contradicts the experimental data. This is explained by the fact that the metal dissolved is carried along by the electron current. It is assumed that almost all valency electrons of the alkali metals are free. With alkali metals, heavy metals form compounds or solvate complexes. These are shifted by the electrons to the anode. An analogous behavior was exhibited by mixtures of K and Na (Fig. 8). With excess K, the solvate complexes consist of Na atoms surrounded by K atoms, and Na diffuses to the anode. With excess Na, the reverse effect appears. The direction of diffusion changes at the point corresponding to the composition of the compound Na_2K . A clear parallelism was found between the coefficient C and $\frac{\partial(\sigma v_{\text{mean}})}{\partial x_2}$, where σ denotes the specific conductivity, v_{mean} the mean atomic volume of the alloy, and x_2 the atomic fraction of K. There are 8 figures and 18 references: 12 Soviet-bloc and 6 non-Soviet-bloc.

ASSOCIATION: Khimiko-tehnologicheskiy institut imeni D. I. Mendeleyeva
(Institute of Chemical Technology imeni D. I. Mendeleyev)

SUBMITTED: September 2, 1959

Card 3/4

Study of electrodiffusion in the alloys...

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S/076/61/035/005/005/008
B101/B218

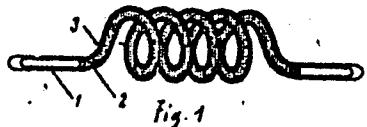
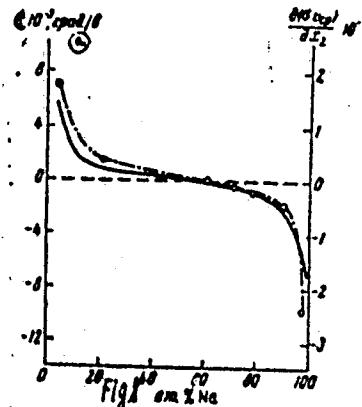


Fig. 1: Spiral tube for measuring electro-diffusion (explanation in the text).

Fig. 8: Coefficient C of electro-diffusion (----) and $\delta(\sigma v_{\text{mean}})/\delta x_2$ (full line) as a function of the composition in the system K - Na.
Legend: a) deg/v



Card 4/4

DRAKIN, S.I.; MIKHAYLOV, V.A.

Calculation of the thermodynamic characteristics of the hydration
of ions incapable of prolonged existence in an aqueous solution.
Zhur.fiz.khim. 36 no.8:1698-1704 Ag '62. (MIRA 15:8)

1. Moskovskiy khimiko-tehnologicheskiy institut imeni Mendeleyeva
1 Institut neorganicheskoy khimii Sibirsogo otdeleniya AN SSSR.
(Ions) (Hydration) (Chemistry, Physical and theoretical)

KAPUSTINSKIY, A.F. [deceased]; MAYYER, A.I.; DRAKIN, S.I.

Heats of solution of potassium chlorides in dioxane-water
mixtures. Trudy MKHTI no.38:10-14 '62. (MIRA 16:7)

(Potassium chloride) (Heat of solution)
(Dioxane)

SERGEYEVA, T.N.; DRAKIN, S.I.

Precise data on stationary distribution during electrodiffusion
in Na - Gd, Na - Pb, and K - Hg alloys. Trudy MKHTI no. 38:103-
108 '62.
(MIRA 16:7)

(Alkali metal alloys—Electric properties)

KUDRYAVTSEV, A.A.; SELIVANOVA, N.M.; DRAKIN, S.I., dots.; MAYYER,
A.I.; SAMPLAVSKAYA, K.K.; SOLOKHIN, V.A.; STAKHANOVA,
M.S.; BUNDEL', A.A., prof., retsentent; KARAPET'YANTS, M.Kh.,
doktor khim. nauk, prof., red.; MEL'NIKOVA, T.I., red.

[Laboratory work in-general and inorganic chemistry] Prakti-
kum po obshchei i neorganicheskoi khimii. [By] A.A.Kudriavtsev
i dr. Moskva, Mosk. khimiko-tehnol. in-t im. D.I.Mendeleeva.
Pt.2. [Work in the chemistry of elements] Raboty po khimii ele-
mentov. 1963. 122 p.
(Chemistry--Laboratory manuals)
(Chemical elements)

(MIRA 16:10)

ACCESSION NR: AR4041542

S/0137/64/000/004/I017/I017

SOURCE: Ref. zh. Metallurgiya, Abs. 41101

AUTHOR: Sergeyeva, T. N.; Drakin, S. I.

TITLE: Temperature dependence of steady-state distribution during electric diffusion in the alloy K-Na

CITED SOURCE: Tr. Mosk. khim.-tekhnol. in-ta im. D. I. Mendeleyeva, vy*p.
41, 1963, 89-90

TOPIC TAGS: steady state distribution, electric diffusion, potassium sodium alloy, potassium alloy, sodium alloy

TRANSLATION: Studies steady-state distribution of electric diffusion in alloy K +3 wt. % Na at 70 and 125°. Steady state was considered attained when the coefficient of electric diffusion, calculated according to method of least squares, remained constant with accuracy of +0.03 degrees/volts for~100 hours. Coefficients of

Card 1/2

ACCESSION NR: AR4041542

electric diffusion calculated by the angle of inclination of straight lines $\log \frac{\Delta p - \Delta E}{\Delta E}$ (Δp is difference between specific electrical resistance of alloy and pure K, ΔE is voltage drop) at 70 and 125° constitute, respectively, $1.28 \cdot 10^4$ and $1.25 \cdot 10^4$ degrees/volts; the difference between these magnitudes is within limits of possible error of the experiment. It is shown that for alloys of an alkali metal it is possible confidently to calculate steady-state distribution in interval $+50^\circ$ by the coefficient of electric diffusion, found at one temperature.

SUB CODE: MM, EM

ENCL:00

Card 2/2

DRAKIN, S.I.

Me - H₂O distances in crystal hydrates and the ionic radii in aqueous solutions. Zhur.strukt.khim. 4 no.4:514-520 Jl-Ag '63.
(NIRA 16:9)

1. Moskovskiy khimiko-tehnologicheskiy institut imeni D.I.Mendeleyeva.
(Ionic crystals) (Water)

S/076/63/037/001/001/029
B101/B186

AUTHORS: Drakin, S. I., Borisova, A. M., Pugatsevich, V. M. (Moscow)

TITLE: Determination of transference numbers on electrodiffusion in Na - Hg, K - Hg, Na - Tl, and K - Tl alloys

PERIODICAL: Zhurnal fizicheskoy khimii, v. 37, no. 1, 1963, 8-12

TEXT: A device for determining the transference number of an alloy that is solid at room temperature and easily affected by air and water vapor is described. 5 g of the alloy is put into a test tube of 17 mm diameter and 20 cm length, air being excluded. Then, a small vessel (0.5 ml) is enclosed in the test tube connected to it through a zigzag tube (diameter 2-2.5 mm, length 12 cm) so as to prevent convection. This inner vessel is filled with the melted alloy by suction. Test tube and inner vessel are provided with electrodes. After the experiment, the composition of the alloy in the test tube and in the inner vessel is analyzed. The concentration of the metal dissolved in the test tube remains almost constant, because the test tube is large compared with the inner vessel. The transference number n is calculated from the difference in content of dissolved metal.

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Determination of transference numbers on ... B101/B186 S/076/63/037/001/001/029

The experiments were conducted at 115 - 215°C, 2.5 - 6 a, 6 - 11 hrs. The values of n as given in a table were used to calculate the diffusion coefficients for Hg dissolved in Na or K and for Tl dissolved in Na or K from the equation $n = DKc\varrho F/T$, where ϱ is the resistivity, T is the absolute temperature, F is the Faraday number, c is the concentration of the metal dissolved, and K is the coefficient of electrodiffusion. The following diffusion coefficients were obtained: for Hg in Na,

$D = 0.70 \cdot 10^{-4} \text{ cm}^2/\text{sec}$ at 115°C; for Hg in K, $D = 1.4 \cdot 10^{-4} \text{ cm}^2/\text{sec}$ at 120°C; for Tl in K, $D = 0.71 \cdot 10^{-4} \text{ cm}^2/\text{sec}$ at 115°C. The thallium atoms become solvated to a higher degree in potassium than in mercury, hence the lower D values for Tl in K. There are 2 figures and 2 tables. The most important English-language reference is: P. Mangelsdorf, J. Chem. Phys., 30, 1170, 1959.

ASSOCIATION: Khimiko-tehnologicheskiy institut im. D. I. Mendeleyeva
(Institute of Chemical Technology imeni D. I. Mendeleyev)

SUBMITTED: December 2, 1960

Card 2/2

DRAKIN, S.I.; FROLOVA, G.M.

Experimental proof of the phenomenological equations describing electro-diffusion in metallic alloys. Zhur.fiz.khim. 37 no.7:1521-1526 J1 '63.
(MIRA 17:2)

1. Moskovskiy khimiko-tehnologicheskiy institut imeni Mendeleyeva.

DRAKIN, Sergey Ivanovich; KUDRYAVTSEV, Aleksandr Andreyevich;
SELIVANOVA, Nadezhda Mikhaylovna; MAYER, Antonina
Ivanovna; SAMPLAVSKAYA, Kira Karlovna; SOLOKHIN, Viktor
Alekseyevich; STAKHANOVA, Mariya Sergeyevna; ALAVERDOV,
Ya.G., red.; FEDOROVA, T.P., red.; KARAPET'YANTS, M.Kh., red.

[Laboratory work in general and inorganic chemistry]
Praktikum po obshchei i neorganicheskoi khimii. Moskva,
Vysshiaia shkola, 1964. 268 p. (MIRA 18:4)

DRAKIN, S.I.; SERGEYEVA, T.N.; TREPAKOV, A.I.

Chemical interaction and electrodiffusion in liquid sodium
alloys. Zhur. fiz. khim. 38 no.2:321-324 F '64.

(MIRA 17:8)

l. Moskovskiy khimiko-tehnologicheskiy institut imeni
D.I. Mendeleyeva, Moskva.

DRAKIN, S.I.; YERBANOVA, L.N.; KARAPET'YANTS, M.Kh. (Moscow)

Determination of instantaneous heat effects by means of the
Mishchenko and Sukhotin modification of the Schottky calorimeter.
Zhur. fiz. khim. 38 no.4:1051-1054 Ap '64. (MIRA 17:6)

l. Moskovskiy khimiko-tehnologicheskiy institut imeni D.I.
Mendeleyeva.

YERBANOVA, L.N.; DRAKIN, S.I.; KARAPET'YANTS, M.Kh.

Comparative study of the heats of solvation of ions in alcohols.
Zhur.fiz.khim. 38 no.11:2670-2674 N '64.

(MIRA 18:2)

1. Moskovskiy khimiko-tehnologicheskiy institut imeni Mendeleyeva.

DRAKIN, S.I.; CHZHAN YU-MIN [Chang Yu-ming]

Comparative study of the heats of solvation of ions in methanol
and water. Zhur. fiz. khim. 38 no.12:2800-2806 D '64.
(MIRA 18:2)
1. Moskovskiy khimiko-tehnologicheskiy institut imeni D.I. Men-
deleyeva.

"APPROVED FOR RELEASE: Friday, July 28, 2000

CIA-RDP86-00513R0004111200

YERBANOV, I.N.; KARPEVICH, M.Kh.; DRAKIN, S.I.

Comparative study of the heat of solvation of ions in alcohols.
Part 2. Zbir.fiz.khim. 39 no.11:2748-2752 N '65.

(MIRA 18:12)

A. N. Moskovskiy khimiko-tehnologicheskiy institut imeni D.I.
Mendelejeva.

APPROVED FOR RELEASE: Friday, July 28, 2000

CIA-RDP86-00513R00041112001

VIL'BERG, S.S. [deceased]; DROZDOV, V.A.; KARATEYEV, D.A. [deceased];
MYSHLYAYEVA, L.V., dots.; SAYUSHKINA, Ye.N.; SENETSKAYA,
L.P.; CHIVIKOVA, A.N.; DRAKIN, S.I., dots., retsenzent

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CA

Cobalt thiocyanate complexes in solution. A. K. Balkin and O. E. Drisko. *Zhur. Obshch. Khim.* (J. Gen. Chem.) 19, 1800-1815 (1949). - Complexes of complexes formed in Me_2CO soln. between Co^{++} and CNS^- were detd. by photometric measurements of the extinction $E = \log I_0/I$, along definite sections across the triangular state diagram $\text{Co}^{++}, \text{CNS}^-, r$ (r = solvent). In Me_2CO , no deviations from Beer's law were found with the mol. ratio $r = \text{Co}^{++}/\text{CNS}^- = 4:1$ to $1:10$, or up to 32-fold dilution, which proves absence of dissociation. (1) At constant total $\text{Co}^{++} + \text{CNS}^- = 0.03 M$, with r varying from $4:1$ to $1:10$, the position of the max. E depends on the wave length of the light. At 650 m μ , $r = 4:1, 1:1, 1:2, 1:4, 1:9$, $E = 0.27, 0.70, 1.10, 1.22, 1.10$; 570 m μ , 0.40, 0.92, 1.22, 1.05, 0.66; 530 m μ , 0.54, 0.70, 0.64, 0.35, 0.24; 500 m μ , 0.55, 0.82, 0.40, 0.31, 0.21. The max. E shifts to longer waves as the concn. of CNS^- is increased. The variation of the position of E max., with the wave length indicates formation of several complexes. II, from the values of E at 530 m μ , the nonnegligible extinction of $\text{Co}(\text{NO}_3)_4$ at the corresponding concns. is deducted. Thus corrected E (i.e., the pure deviations from additivity) are (for the above r) 0.37, 0.58, 0.88, 0.30, 0.21, with a max. at $1:1$, i.e., at the compn. $\text{Co}(\text{CNS})_4$. The diminution of E with further decreasing r (increasing CNS^-) indicates gradual disappearance of that complex, and formation of complexes with greater mol. of CNS^- . The complex $\text{Co}(\text{CNS})_4$ dets. the max. in 670 m μ , situated at $r = 1:2$, and the blue complex $\text{Co}(\text{CNS})_4^{++}$ dets. the max. in 500 m μ . (2) These results are confirmed by similar determinations at const. concn. of $\text{Co}^{++} = 0.01 M$, with r varying from $10:1$ to $1:20$. In 500-530 m μ , decrease of r to $1:1$ and $1:2$ deepens the color; further decrease weakens it. This indicates disappearance of $\text{Co}(\text{CNS})_4$ and $\text{Co}(\text{CNS})_4^{++}$, and formation of $\text{Co}(\text{CNS})_4^{--}$ and $\text{Co}(\text{CNS})_4^{---}$, which absorb in the red. Further increase of the excess of CNS^- causes no change; consequently no complexes are formed higher than $\text{Co}(\text{CNS})_4^{---}$, and the latter complex requiring no excess CNS^- , has evidently no tendency to dissolve in Me_2CO . (3) The existence of intermediate complexes

appears particularly clearly in series with a const. concn. of $\text{CNS}^- = 0.01 M$, and variable concn. of Co^{++} , r from $1:6$ to $4:1$. In 650 m μ the color intensity first increases with increasing Co^{++} ; further increase, beyond the compn. $\text{Co}(\text{CNS})_4$, weakens the color, owing to the reactions $3\text{Co}(\text{CNS})_4^{--} + \text{Co}^{++} \rightarrow 4\text{Co}(\text{CNS})_4^{--}$, and $2\text{Co}(\text{CNS})_4^{--} + \text{Co}^{++} \rightarrow 4\text{Co}(\text{CNS})_4^{--}$. In 500 m μ , E (corrected for the proper extinction of $\text{Co}(\text{NO}_3)_4$) increases up to $r = 1:1$, then weakens with further increasing Co^{++} , owing no doubt to further interaction between $\text{Co}(\text{CNS})_4^{--}$ and $\text{Co}(\text{NO}_3)_4$, resulting in complexes still poorer in CNS^- . (4) The existence of such complexes is demonstrated by the spectroscopic methodically ascertained nonadditivity between $\text{Co}(\text{CNS})_4^{--}$ and $\text{Co}(\text{NO}_3)_4$. The extinction curve of a soln. $r = 2:1$ is not intermediate between the curves of $\text{Co}(\text{CNS})_4^{--}$ and $\text{Co}(\text{NO}_3)_4$, but shows a max. at 580 m μ . This is most probably due to a complex of $\text{Co}(\text{CNS})_4^{--}$ with the solvent. In an acid soln. (HClO_4 , 0.1 M) additivity is preserved, the position of the max. for $r = 2:1$ being the same (580 m μ) as for 1:1. The intensity of the color is lowered, owing to a lowering of the dissociation of HCN by HClO_4 . (5) Formation of the blue $\text{Co}(\text{CNS})_4^{++}$ ion is further confirmed by transference data. N. Plotow

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A.K. Babko and O.F. Drako. Colorimetric determination of cobalt as a rhodanide complex.
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CARDFK-0, O. F.

Effect of the solvent on the dissociation of the cobalt thiocyanate complex. A. K. Babko and O. F. Drakova (Acad. Sci., Ukr. SSR) Zhur. Obshch. Khim. (J. Gen. Chem.) 20, 228-34 (1950); cf. C.A. 44, 1355A.

The equil. const. $K = [B][A]/[BA]$ of the colored complex BA formed between the central cation B and the anion B is evaluated from the excess of [A] necessary for the optical d. D of the soln. to attain half the max. D_{max} corresponding to a very large excess of A; at

$D = 1/2 D_{max}$, $[BA] \approx [B]$, and $K \approx [A]$. By photocalorimetry of solns. of $\text{Co}(\text{NO}_3)_2 + \text{KCNS}$, in 610 m μ , in mixts. of $\text{H}_2\text{O} + \text{Me}_2\text{CO}$ (vol. % H_2O) = $-\log K$ (for the complex ion $[\text{Co}(\text{CNS})_4]^{2-}$) = 2.4, 2.0, 1.6, 1.15, and -0.6, resp.; the dielectric constants of these solns., calcd. by the additivity rule, are 27, 38, 51, 60, and 81, resp. In the range 10-75 vol. % H_2O , there is a near proportionality between $-\log K$ and 1/ ϵ . The same holds approx. for mixts. of H_2O with AcOH , dioxane, EtOH , and HCO_2H . However, there is no such proportionality from one solvent to another; examples of data of $-\log K$ (in parentheses, 1/ ϵ) are: $\text{AcOH} + 10, 25, 50, 75$ vol. % H_2O , 2.1 (0.071), 1.7 (0.043), 1.4 (0.021), 1.3 (0.016); dioxane + 10, 25, 75, 100 vol. % H_2O , 2.2 (0.100), 1.8 (0.048), 1.6 (0.024), 1.4 (0.018); $\text{EtOH} + 10, 25, 50, 75$ vol. % H_2O , 2.1 (0.031), 1.7 (0.026), 1.6 (0.019), 1.1 (0.015); $\text{HCO}_2\text{H} + 10, 25, 50, 75$ vol. % H_2O , 1.7 (0.020), 1.4 (0.019), 1.0 (0.018), 0.8 (0.014). If $-\log K$ is plotted against the $[\text{H}_2\text{O}]$ content of the solvent, the curves for Me_2CO , AcOH , dioxane, EtOH , and HCO_2H are of the same shape and lie closely together. This similarity indicates a common mechanism, whereby the nonaq. component of the solvent promotes removal of H_2O from the inner sphere of the complex and thus facilitates complex formation according to $[\text{Co}(\text{H}_2\text{O})]^{2+} + 4\text{CNS}^- \rightarrow [\text{Co}(\text{CNS})_4]^{2-} + 4\text{H}_2\text{O}$. This accounts for the absence of any marked specificity of the nonaq. solvent.

N. Thom

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PAGE 1 BOOK EXTRICATION 007/2433

DR-AKO-QF
Aksel'yev and S.S. Kostyleva po smeshcheniyu kisli
metally opredeleniya prinyaty v chistym metallicheskikh metodakh po determinatsii analiz
tsev v rukovodstve

Novyj opredeleniya prinyaty v chistym metallicheskikh metodakh po determinatsii analiz
tsev v rukovodstve. Moscow, 1960. 111 p. (Series: Issled. Tsv. Tsvet. 12) 5,500
kopij peredatek.

Byap. Rar.: I.P. Vaynshteyn, Redaktsion. i. D.I. Rybachikov. Doctor of Chemical
Sciences; Ed. of Publishing House: M.P. Tolpugina. Ed.: T.V. Poljakova.

PURPOSE: This collection of articles is intended for chemists, metallurgists, and
engineers.

CONTENTS: The articles describe methods for detecting and determining various ad-
mixtures and their traces in pure metals, also discuss many analytical, physicochemical, electrochemical, spectrophotometric and luminescent methods of
analyzing materials of high purity. The editor notes that these methods have
been developed within the last time, or else from factory laboratories of the
authorities, and are now widely used in research and factory laboratories of the
various plants. No personalities are mentioned. References: mostly Soviet.

Accompany each article.

Editor: A.Y. Drabkin. Institute of Determinative Analysis in Metallurgy (Moscow) 202

Spectrophotometric Methods of Determination of Admixtures in Metallic Compounds and
Oxides in Metals

Babkin, A.L. and V.P. Poljakov. Spectroscopic Detection of Small Quantities
of Oxygen in Metallic Compounds

Babkin, A.L., V.P. Poljakov, and O.P. Drabkin. Determination of Small Quantities
of Oxygen in Metallic Compounds

Babkin, A.L., V.P. Poljakov, and O.P. Drabkin. Determination of Tantalum
in Tantalum in the Ferromanganese

Babkin, A.L., V.P. Poljakov, and I.A. Romanova. Determination of Ad-
mixtures of Lead, Tin, Copper, Zinc, and Cobalt in Tantalum and in Tungsten

Allova, 72

Babkin, A.L. Spectrographic Determination of Manganese and Titanium in
Oxides and Minerals

Babkin, A.L., V.P. Poljakov, I.V. Kostyleva, M.P. Tolpugina, V.V.
Korshak, and Yu. I. Kostyleva. Spectrophotometric Method of Determination of Ad-
mixtures of Lead, Tin, Copper, Zinc, and Cobalt in Manganese, Nickel,
and Titanium

Babkin, A.L. Spectrographic Determination of Manganese and O.F. Diamant.
Determination of Metallic Inclusions in Titanium and Zirconium

Babkin, A.L., Yu. I. Kostyleva, V.P. Poljakov, and S.E. Sazanova.

Determination of Admixtures in Titanium and Zirconium Oxides

Babkin, A.L., Yu. I. Kostyleva, V.P. Poljakov, and S.E. Sazanova. Determination of Metallic Inclusions
in Zirconium Oxide. Optical Method

Babkin, A.L., Yu. I. Kostyleva, V.P. Poljakov, and S.E. Sazanova. Determination of the Percentage of
Oxygen in Zirconium From the Content of Unconverted Ox. Phases at Various
Growth Temperatures

Babkin, A.L., Yu. I. Kostyleva, V.P. Poljakov, and S.E. Sazanova. Determination of Oxygen in Titanium
and in Zirconium by the Vacuum-Fusion Method

Babkin, A.L., Yu. I. Kostyleva, V.P. Poljakov, and S.E. Sazanova. Determination of Small Quantities of
Oxygen in Zirconium

Babkin, A.L., Yu. I. Kostyleva, V.P. Poljakov, and S.E. Sazanova. Determination of the Percentage of
Oxygen in Zirconium From the Content of Unconverted Ox. Phases at Various
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BABKO, A.K.; VOLKOVA, A.I.; DRAKO, O.P.

Determining small amounts of oxygen in metallic germanium. Trudy Kom.
(MIRA 13:8)
anal. khim. 12:52-64 '60.
(Germanium--Analysis) (Oxygen--Analysis)

CECH, Oldrich; JAKOVA, Sona

Studies on the state of physical development of working adolescents
in Slovakia. Pracovni lek. 11 no.6:299-304 Aug 59.

1. Vyzkumny ustav telovychovny, red. MUDr. Merhautova Ustav hygiény
prace a chorob z povolani, red. prof. Teisinger.
(ADOLESCENCE)

BENA, E.; PODLESAK, K.; JOKL, M.; DRAKOVA, S.

Work physiology in refrigeration plants. Pracovni lek.12 no.10:
526-533 D '60.

(REFRIGERATION)
(OCCUPATIONS AND PROFESSIONS)

DRAKOWA, Danuta; MICHEJDA, Adam.

Difficulties in the diagnosis of abdominal symptoms in the course of Schonlein-Henoch disease. Pediat. polska 30 no.12: 1177-1184 Dec 55.

1. Z I Kliniki Pediatricznej A.M. we Wrocławiu. Kierownik: prof. dr. med. H.Hirschfeldowa i z II Kliniki Chirurgicznej A.M. we Wrocławiu. Kierownik: prof. dr. med. W.Bross. Wrocław, ul. Nulla 9 m. 1.

(PURPURA, NORMTHROMBOPENIC

Schonlein-Henoch purpura, abdom. manifest., diag.)

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Schonlein-Henoch purpura)

HIRSZFELDOWA, Hanna; DRAKOWA, Danuta; PODOLAK, Olga

Antibodies in maternal milk as a cause of infantile hemolytic jaundice. Pediat.polska 35 no.12:1401-1406 D '60.

1. Z I Kliniki Pediatricznej A.M. we Wrocławiu, Kierownik: prof.
dr med. H.Hirschfeldowa.
(JAUNDICE HEMOLYTIC in inf & child)
(MILK HUMAN)

NOWAKOWSKI, T.K.; CHABUDZINSKA, S.; DADEJOWA, J.; DRAKOWA, D.; SIENNICKI, W.

Evaluation of the agglutination of *Corynebacterium diphtheriae*.
Pediat. polska 36 no.4:369-374 '61.

l. Z I Kliniki Pediatricznej Kierownik: prof. dr H. Hirsfeldowa
Z Kliniki Chorob Zakaznych Wieku Dziecięcego Kierownik: prof. dr
T.K. Nowakowski Z Wojewódzkiej Stacji Sanitarno-Epidemiologicznej
Kierownik: dr St. Przylecki.

(CORYNEBACTERIUM DIPHTHERIAE)

POLAND

NOWAKOWSKI, Tadeusz K., DRAKOWA, Danuta, DADEJOWA, Janina, and HAJZIK, Roman, Clinic for Infectious Diseases of Children (Klinika Chorob Zakaznych Wieku Dziecięcego), Medical Academy (Akademia Medyczna) in Wrocław (Director: Prof. Dr. med. T. K. NOWAKOWSKI)

"Possibility of Limiting Infections Inside Hospital."
Warsaw, Przeglad Epidemiologiczny, Vol 17, No 3, 63, pp
253-255

Abstract: [Authors' English summary] The hospital with which the authors are affiliated is located in a typical hospital building. After slight adaptations of the building to achieve good work organization, the number of infections in the hospital dropped to 0.6%, despite the fact that during that time 642 children with 17 different infectious diseases were treated there, 263 of which had viral infections. There are no references.

1/1

KOSTOLOWSKI, Roman; DRAKOWA, Danuta; NOWAKOWSKI, Tadeusz K.; HAJZIK, Roman

Simulated and true otitis media in measles in young children.
Otolaryng. pol. 17 no.4:495-496 '63.

1. Z Kliniki Chorob Zakaznych Wieku Dziecięcego AM we Wrocławiu. Kierownik: prof. dr. T.K.Nowakowski.

POLAND

DRAKOWA, Danuta; Clinic of Child Infectious Diseases (Klinika Chorob Zakaznych Wieku Dzieciecego), AM [Akademia Medyczna, Medical Academy] in Wroclaw (Director: Prof. Dr. Tadeusz NOWAKOWSKI)

"Toxic-Allergic Syndrome After Too Large Doses of Retasulfine in Children. Report of Four Cases."

Warsaw, Polski Tygodnik Lekarski, Vol 18, No 15, 8 Apr 63, pp 547-550.

Abstract: [Author's English summary] The author reports four cases of toxic-allergic syndrome after too large doses of retasulfine in children, and draws attention to the hemorrhagic character of the rash, similar to that in morbilli or scarlet fever. All cases exhibited conjunctivitis, enlargement of liver and nodes, prolongation of clotting time, and leukopenia with fever, and sporadically eosinophilic jaundice and cerebral signs were noted. Some of the signs have not previously been reported in the literature. There are 12 references, of which two (2) are in Polish, three (3) each in French and German, and four (4) in English.

T/1

JEZIORO, Zdzislaw; DRAK, Juliusz; DRAKOWA, Danuta

Apropos of the treatment of cardiospasm in children. Pol. tyg.
lek. 19 no.21:799-800 18 My'64

l. Z III Kliniki Chirurgicznej Akademii Medycznej we Wrocławiu
(kierownik: prof. dr. med. Z. Jeziorko) i z I Kliniki Chorób
Dziecięcych Akademii Medycznej we Wrocławiu (kierownik: prof. dr.
med. T. Nowakowski).

DRAK, Juliusz; DRAKOWA, Danuta; GOLEN-TETER, Maria

A case of perforated intestinal cyst in a child. Pediat. Pol.
39 no.7:845-847 Je '64.

1. Z III Kliniki Chirurgicznej AM we Wrocławiu (Kierownik:
prof. dr med. Z. Jezioro) i z I Kliniki Pediatricznej Akademii
Medycznej we Wrocławiu (Kierownik: prof. dr med. T.
Nowakowski).

DRAKULIC, Dusan, ing. (Beograd, Uralska 22)

Analysis of the sample distribution network of the sectionally impregnated ore body of the lead-zinc deposits of Kisnica Mine. Tehnika Jug 17 no.4:673-679 Ap '62.

1. Chief of the Geological Division of the Iron Ore Exploration and Exploitation Enterprise under Construction, Titovo Uzice.

DRAKULIC, Dusan, inz. (Beograd, Neznanog junaka 19); DEDIC, Ljubomir, inzh.,
geolog.

Structural composition of the Upper Cretaceous sediments of
Mokra Gora and Beli Rzav. Tehnika Jug 18 no.10:Supplement:
Rudarstvo metalur 14 no.10:1858-1864 063.

1. Sef geoloske službe Preduzeća za istraživanje i eksploataciju gvozdene rude u izgradnji, Titovo Uzice (for Drakulic).
2. Preduzeće za istraživanje i eksploataciju gvozdene rude u izgradnji, Titovo Uzice (for Dedic).

DRAKULIC, Dusan (Beograd, Neznanog junaka 19)

Extent of prospecting work and classification of ore reserves in the lead and zinc deposits of Ajvalija, Prline, and Kisnica. Tehnika Jug '8 no. 12: Supplement: Rudarstvo metalurg 14 no. 12: 2237-2243 D '63

1. Sef geoloske sluzbe Preduzeca za istrazivanje i eksploataciju gvozdene rude u izgradnji, Titovo Uzice.

DRAKULIC, Dusan, inz. (Beograd, Nenzznanog Junaka 19)

Outlining the vertically impregnated Pb-Zn deposit at Kisnica.
Tehnika Jug 19 no.1:Suppl;Rudarstvo metalurg 15 no.1:82-83 Ja '64.

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gvozdene rude u izgradnji, Titovo Uzice.

DRAKULIC, Krunoslava, dr

Indications for otorhinolaryngological surgery in aviation. Voj.san.
pregl., Beogr. 17 no.12:1309-1311 D '60.

1. Vuzduhoplovnomedicinski institut u Zemunu
(AVIATION MEDICINE)
(OTORHINOLARYNGOLOGY surg)

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DRAKULIC, Krunoslava, Vojni sluzbenik II klase, dr.

Air disease from an otological viewpoint. Voj.san.pregl. 18 no.2:
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Changes of the respiratory mucous membranes under the influence of
pure oxygen. Vojnosanit. pregl. 19 no.11 774-776 N '62.

1. Vazduhoplovnomedicinski Institut u Zemunu.
(RESPIRATORY SYSTEM) (OXYGEN)

RISAVI, Antun, sanitetski pukovnik docent dr.; DRAMULIC, Krunoslava, vojni sluzbenik I klase dr.; MIHAJLOVIC, Radivoje, sanitetski tehnicar

The role of speech audiometry in the evaluation of hearing disorders in pilots. Vojnosanit. pregl. 22 no.10:616-620 O '65.

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DRAKULIC, Marija, major mr. ph.; BINNIFELD, Zlatko, major mr. ph.

Neurotoxic war gases. Voj. san pregl., Beogr. 11 no.9-10:
378-384 Sept-Oct 54.

(GASES

war gases, neurotoxic, chem. & physiol. eff.)

SMIT, Slavica; MILETIC, Branimir; DRAKULIC, Marija; STAVRIC, Stanislava;
BRDAR, Branko

Photorestitution of the biosynthesis of nucleic acids in
irradiated bacteria. Biol glas 15 no. 4: 207-214 '62.

1. Institut" Ruder Boskovic", Radiobioloski odjel, Zagreb.

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Incorporation of adenine-14C into nucleic acids of bacteria
irradiated with ultraviolet rays. Biol glas 15 no. 4: 215-
223 '62.

1. Institut "Ruder Boskovic", Radiobiolski odjel, Zagreb.

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Restoration of irradiated animal cells with isologous subcellular fractions. Vojnosanit Pregl. 20 no.10:629-635 O '63.

1. Institut "Ruder Boskovic," Zagreb, radiobioloski odjel.

KOS. E.; DRAKULIC, M.

Possible preservation of desoxyribonucleic acid integrity
following "in vivo" irradiation with ionizing radiations. II.
Inhibition of desoxyribonucleic acid degradation with gamma-
rays with the aid of some Kræta' cycle metabolites.
Vojnosanit Pregl. 21 no.2:89-91 F '64.

1. Institut "Ruder Boskovic", Zagreb, radiobioloski odjel
i Poljoprivredni fakultet Sveucilista u Zagrebu, kemijski
zavod.

DRAKULIC, M.; KOG, E.

A possibility of the preservation of the value of deoxyribonucleic acid after ionizing irradiation in vivo. III. Inhibition of produced by gamma-rays of deoxyribonucleic acid breakdown due to some metabolites participating in energy metabolism and some of their antagonists. Vojnosanit pregl. 21 no.3:169-171 Mr '64.

1. Radiobiolski odjosi, Institut "Ruder Boskovic" u Zagrebu.

SMIT, S.; MILETIC, B.; GIGOV, A.; BOGDANOVIC, M.; DANON, J.; JANKOVIC, M.M.;
CIPINA, T.; MILOSEVIC, R.; JANKOVIC, M-a; BOGCEVIC, R.; STAVRIC, S.;
DRAKULIC, M.; MATONICKIN, I.; PAVLETIC, Z.

Review of periodicals; biology. Bul sc Youg 9 no.4/5:138-
139 Ag-0 '64.

DRAKULIC, P.

Graphite carbon as antifrictional material in machinery. p. 1545.

TEHNIKA. Beograd, Yugoslavia. Vol. 14, no. 9, Sept. 1959.

Monthly List of East European Accessions (EEAI) LC Vol. 9, no. 2, Feb. 1960.

Uncl.

Y/010/62/000/003/002/002
I035/I235

AUTHOR: Dražilic, Pavle, Engineer

TITLE: Principles of construction of rockets with liquid propellants

PERIODICAL: Vojnotehnický Glasnik, no. 3, 1962, 200-208

TEXT: Popular description of the main components of liquid-fuel rocket engines. Different types of combustion chambers (cylindrical, conical, spherical), injectors (stream, centrifugal), fuel tanks and other accessories are mentioned; valves, pumps and other components are not discussed. Schematic drawings of many components are given. No mention is made of any technical data. There are 15 figures.

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DRAKULIC, T.

New materials for the manufacture of furniture. p. 303.

NARODNI SMAR. (Drustvo sumarskih inzenjera i tehnicara Bosne i Hercegovine)
Sarajevo, Yugoslavia. Vol. 13, no. 5/6, 1959.

Monthly List of East European Accessions (EEAI) ^{L2} Vol. 9, no. 2, Feb. 1969.

Uncl.

L 37215-66 EWP(j)/EWT(m)/T IJP(c) RM/WW

2

ACC NR: AP6018127 (A) SOURCE CODE: UR/0191/66/000/006/0040/0042

AUTHOR: Koyarskaya, B. M.; Kolesnikov, G. S.; Levantovskaya, I. I.; Smirnova, O. V.; Drakyuk, G. V.; Poletakhina, L S.; Korovina, Ye. V.

41
B

ORG: none

TITLE: Thermo-oxidative degradation of polycarbonates

SOURCE: Plastichekiye massy, no. 6, 1966, 40-42

TOPIC TAGS: polycarbonate plastic, heat resistance, oxidative degradation, oxidation kinetics, reaction mechanism

ABSTRACT: Polycarbonates, molecular weight of about 30,000, based on 2,2-di-(4-hydroxyphenyl)-propane (PK-1), on 1,1-di-(4-hydroxyphenyl)-cyclohexane (PK-2) and on di-(4-hydroxyphenyl)-phenylmethane (PK-3) were subjected to thermal oxidation in vacuum.¹ Kinetic curves of the thermal oxidations showed PK-1 was most stable and PK-3 the least stable. Energies of activation for the oxidations were calculated: 21.0, 17.6 and 13.0 kcal/mol, respectively. Reaction mechanisms are discussed. Auto-accelerated processes are indicated in the initial period of thermal oxidation of PK-1 and PK-2. Radical-chain oxidation

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L 37215-66

ACC NR: AP6018127

mechanisms are indicated for all three materials. It is concluded that polycarbonates with increased resistance to thermal oxidation should contain a minimum number of "aliphatic" hydrogen atoms in the main polymer chain and the bisphenols with aliphatic carbon atoms bonded to hydrogen. Orig. art. has: 3 figures, 6 equations and 2 formulas.

SUB CODE: 07/ SUBM DATE: none/ ORIG REF: 009/ OTH REF: 002

ns
Card 2/2

VECHTOMOV, M.I., inzh.; KUDRYAVTSEV, V.A., inzh.; MALKES, D.A., inzh.; OSTROVSKIY, G.I.; POVERENNYY, L.D.; SUSHKOV, P.M., inzh.; TYULENEV, N.Z., inzh. Prinimali uchastiye: GALYAMOVA, N.S., inzh.; PUTEYEVA, N.P.; IZRAYLOVICH, Ye.A., inzh.; MARCHENKO, G.A., inzh.; MALYGINA, Z.S.; SOKOLOVA, Ye.A.; SOKOV, V.N., inzh.; TARASOVA, S.N.; TASHAYEV, A.L., inzh.; FILIMONOV, S.V.; DRALICH, K.F., inzh., nauch. red.; NOVITCHENKO, K.M., inzh., nauchnyy red.; SIMAKOV, S.N., inzh., nauchnyy red.; FAKTOROVICH, Yu.A., kand. tekhn. nauk, nauchnyy red.; STUPIN, Ye.N., otv. red.; LUTOV, N.S., red.; IVANOV, V.S., red.; BAGUZOV, N.P., glav. red.; VOLCHEGORSKIY, M.S., zam. glav. red.; DOBRYNIN, S.N., red.; NAZAROV, I.A., red.; KOLESNIKOV, S.I., red.; MEL'NIKOV, N.P., red.; SUSNIKOV, A.A., red.; STAROVEROV, I.G., red.; LYTKINA, L.S., red. izd-va; GORDEYEV, P.A., red. izd-va; OSENKO, L.M., tekhn. red.

[Handbook for the designer of industrial, residential, and public buildings and structures; organization of construction and execution of building and assembly operations. Industrial construction] Spravochnik proektirovshchika promyshlennykh, zhilykh i obshchestvennykh zdanii i sooruzhenii; organizatsiya stroitel'stva i proizvodstvo stroitel'no-montazhnykh rabot. Promyshlennoe stroitel'stvo. Pod red. P.M. Sushkova. Moskva, Gos.izd-vo lit-ry po stroit., arkhit. i stroit. materialam, 1961. 372 p.

(MIRA 15:2)

(Industrial buildings)

NAZAROV, Vasiliy Stratonikovich; MUROMTSEV, Aleksey Mikhaylovich; DZAL-
XIN A.G., redaktor; KAN, P.M., redaktor; KRASHAYA, A.K., tekhnicheskiy redaktor.

[Oceanography] Okeanografiia. Moskva, Izd-vo "Morskoi transport,"
1954. 165 p.
(Oceanography) (MLRA 7:12)

DRALKIN, A.G.

Drift station "North Pole 4" in the polar region of the Central
Polar Basin. Mor. flot 18 no. 6:21-22 Ja '58. (MIRA 11:7)

1. Nachal'nik otdela nauki i gidrografii Gidrograficheskogo
upravleniya Glavsevmorputi.
(Arctic regions)

DRALKIN, A.G.

Radio messages from Antarctica. Inform. biul. Sov. antark. eksp.
no. 9:48-50 '59
(MIRA 13:3)

1. Nachal'nik Chetvertoy kontinental'noy ekspeditsii.
(Antarctic regions)

DRALKIN, A.G.

Radio messages from Antarctica. Inform. biul. Sov. antark. eksp.
no.10:34-35 '59 (MIRA 13:3)

1. Nachal'nik Chetvertoj kontinental'noy ekspeditsii.
(Antarctic regions)

DRAKIN, A.G.

Radio broadcasts from Antarctica. Inform. biul. Sov. antark. eksp.
no. 11:46-48 '59. (MIRA 13:5)

1. Nachal'nik Chetvertoy kontinental'noy ekspeditsii.
(Antarctic regions)

84583

9,9100 (2101,1041,1046)

S/169/60/000/009/001/007
A005/A001

Translation from: Referativnyy zhurnal, Geofizika, 1960, No. 9, pp. 1-2, #
10065

AUTHOR: Dralkin, A.G.

TITLE: The Mirnyy Observatory. - The Vostok Station. - The Lazarev
Station

PERIODICAL: Inform. byul. Scv. antarkt. ekspeditsii, 1959, No. 12, pp. 47-49

TEXT: The author reports on the works carried out at the Mirnyy Observatory and the stations Vostok and Lazarev during June 1959. The temperature of air at the Mirnyy Observatory during June ranged from -3.8°C to -27.4°C ; the average value of the atmospheric pressure was 987.2 mb, that of the wind velocity was 14.4 m/sec. From June 20 to 30, the snow transfer amounted to 56,468 t along the coast line 1 km long and 3.5 m high. At the Vostok Station, the temperature of air during the month ranged from -53.7°C to -77.6°C . The average value of the atmospheric pressure was 623.8 mb, that of the wind velocity was 4.9 m/sec. The minimum temperature of air (-84.7°C) was recorded

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A005/A001

The Mirnyy Observatory. - The Vostok Station. - The Lazarev Station

at the altitude of 18,040 m. In the Indian sector of the Antarctic, the zonal circulation predominated. In the Mirnyy-Station region, low atmospheric pressure was recorded at the beginning of June and during the time from June 10 to 20, which was caused by filling up of cyclones. The wind of the eastern quarter predominated in the troposphere. In the lower stratosphere, jet streams of the western quarter emerged periodically. In the lower troposphere, pressure inversions were stated, the vertical thickness of which ranged in the limits of 0.5-1 km. The magnetic field was comparatively quiet during a month. A weak storm was noted on June 11. The Ionosphere was in a quiet state. According to the observations of the Vostok Station, the ionosphere was characterized by some decrease in the diffusivity of the F layer and approximately constant presence of the E layer. At the end of June, cases of the appearance of the sporadic screening E layer became considerably more frequent. The Mirnyy Observatory began observations of the temperature state of the shore ice; the Vostok Station carried out three snow surveys in the open section. From June 23 to 24, a collapse of the barrier occurred eastward of the Khmara

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S/169/60/000/009/001/007
A005/A001

The Mirnyy Observatory. - The Vostok Station. - The Lazarev Station

Cape, in consequence of which a big amount of ice fell upon the sea and several
icebergs were formed. The collapse destroyed partially the shore ice, hummocks
occurred. Blocks of bottom ice appeared at the surface in the zone of the col-
lapse.

N.V. Zolotavina

Translator's note: This is the full translation of the original Russian ab-
stract.

Card 3/3

DRALKIN, Aleksandr Gavrilovich, kand. geograf.nauk; FAYNBOIM, I.B.,
red.; SAVCHENKO, Ts.Vr, tekhn.red.

[Recent explorations in Antarctica; work results of the Fourth
Soviet Antarctic Expedition] Novye issledovaniia v Antarktike;
k itogam raboty 4-i Sovetskoi antarkticheskoi ekspeditsii.
Moskva, Izd-vo "Znanie," 1960. 47 p. (Vsesoiuznoe obshchestvo po
rasprostraneniiu politicheskikh i nauchnykh znanii. Ser.9,
Fizika i khimiia, no.21). (MIRA 14:1)

1. Nachal'nik 4-y Sovetskoy antarkticheskoy ekspeditsii (for
Dralkin).
(Antarctic regions--Russian exploration)

DRALKIN, A.G.

Preliminary work results of the Fourth Continental Antarctic
Expedition. Probl.Arkt.i Antarkt. no.5:7-14 '60. (MIRA 14:4)

(Antarctic regions -Russian exploration)

DRAIKIN, A.G.

Radio broadcasts from Antarctica. Inform.biul.Sov.antark.eksp.
no.14:33-35 '60. (MIRA 13:6)

1. Nachal'nik Chetvertoy kontinental'noy ekspeditsii.
(Antarctic regions—Russian exploration)